

## Guided Reading Chapter 13 Section 2

- All compounds have an electrical charge of \_\_\_\_\_.
  - one
  - two
  - zero
  - ten
- An oxidation number is the quantity that indicates the charge on an atom when it has gained, lost or \_\_\_\_\_ electrons.
- Copy figure 13.12, showing oxidation numbers of some common elements.
- Would Beryllium tend to lose two electrons or gain six when forming bonds?
- What is the most common oxidation number for group three on the Periodic table?
- Elements near the noble gases tend to form \_\_\_\_\_ bonds.
  - ionic
  - covalent
  - metallic
- The farther apart elements are on the Periodic Table the more likely they are to form \_\_\_\_\_ bonds.
  - ionic
  - covalent
  - metallic
- Nonmetals tend to form \_\_\_\_\_ bonds.
  - ionic
  - covalent
  - metallic
- Using figure 13.14, how many Chlorine atoms are needed to bond with a Copper (II) atom to form a compound?
- What is a binary compound?
- How do you write the name of a binary compound if given the chemical formula?

12. How many atoms of each element is in  $\text{CaCO}_3$ ?

13. What type of ion is one that contains more than one atom?

14. What is the oxidation number for peroxide?

15. How do you name a polyatomic ion if given its chemical formula?